

Warm up:



Study Notes/Questions

Chemical Formula

- shorthand method of _____
- _____ of the component elements
- Total ion charge = _____

Binary Ionic Compounds

- ionic compounds with _____
- Positive ion (metal ion) _____, negative ion _____
- Use _____ to show numbers of each element needed

Formulas for Binary Ionic Compounds

1. Need symbol and ion charge
2. Balance ion charges to complete formula.

E.g.

1) Calcium combining with oxygen

2) potassium with nitrogen

3) Aluminum with sulfur

Binary Compounds with Multivalent Elements

_____ – can form ions with more than one charge (transition metals)

You will be told which one to use in each situation

E.g. copper can be 1+ or 2+

manganese (III) and sulfur

7.3 Part 1 - Ionic Compounds: Chemical Formulas and Naming

Study Notes/Questions

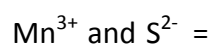
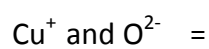
Ionic Compounds: Naming

The chemical names are derived _____.

Remember non-metal ion names _____.

For example:

For multivalent metals use _____ to show the charge on the ion



Summary: (two to three sentences summarizing this section)

**Self-Reflection
Questions:**

1. Describe one thing you learned about this topic today.

2. Describe one thing about this topic you want to learn in more detail.