

Warm up:



Study Notes/Questions

**Classifying Inorganic Compounds**

Compounds with a high % of carbon (by mass) are \_\_\_\_\_

Others are \_\_\_\_\_

Most compounds studied so far are \_\_\_\_\_

Inorganic compounds – \_\_\_\_\_

\_\_\_\_\_ - covalent bonds, non-metal elements  
- small group, but very common

E.g.

\_\_\_\_\_ – ionic bonds, metal + non-metal  
- larger category

E.g.

Inorganic ionic compounds can be \_\_\_\_\_  
depends on properties and ions present

ACID – releases \_\_\_\_\_ in solution

BASE – releases \_\_\_\_\_ in solution

SALT – releases \_\_\_\_\_ in solution, but \_\_\_\_\_

Symbol: (aq) =

**Indicators**

Molecular compounds with a \_\_\_\_\_

React with acid and/or base to form \_\_\_\_\_  
with a different colour

## 8.1 Classifying Inorganic Compounds

Study Notes/Questions

### Acidity

measures the relative amounts of \_\_\_\_\_ in a solution

If  $H^+$  and  $OH^-$  are equal, solution is \_\_\_\_\_

If  $H^+ > OH^-$ , solution is \_\_\_\_\_

If  $OH^- > H^+$ , solution is \_\_\_\_\_

often measured on a \_\_\_\_\_

### pH

The \_\_\_\_\_ the solution, the \_\_\_\_\_ the pH  
Each number (below 7) is 10 times more acidic

The \_\_\_\_\_ the solution, the \_\_\_\_\_ the pH  
Each number (above 7) is 10 times more basic

Example:

pH 4 is \_\_\_\_\_ than pH 5

pH 3 is \_\_\_\_\_ than pH 5

pH 7 is \_\_\_\_\_, it is neutral

Summary: (two to three sentences summarizing this section)

**Self-Reflection  
Questions:**

1. Describe one thing you learned about this topic today.

2. Describe one thing about this topic you want to learn in more detail.